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3. Pure substances

refer to the

substances which

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have an unchanging chemical composition all over them. For example, Water, gold, Pure metals are some of pure metals.

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container has volume

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of, and holds steam
at C. $1/4$ of the
volume is in liquid
point and the
remaining at vapor
form. Determine the
pressure of the steam,
and quality of the
saturated mixture,
and density of the
mixture. Given:

Volume (V)

Temperature (T) C.

Find: The pressure of

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Substances Phase

Change and Property

Diagrams In this

chapter we consider

the property values

and relationships of a

pure substance (such

as water) which can

exist in three phases

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Solid, liquid and
gas.

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Substances –
Thermodynamics

The change in
internal energy can
be found from the
first law of

thermodynamics: U

$= Q - W = (3.5 \times 10$

$\text{J}) - (2.1 \times 10 \text{ J}) =$

$0.9 \times 10 \text{ J} = 90 \text{ kJ.}$)

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A gas in a cylinder is kept at a constant pressure of 3.5×10^5 Pa while 300 kJ of heat are added to it, causing the gas to expand from 0.9 m^3 to 1.5 m^3 .

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The missing thermodynamic properties of water are to be determined using EES, and the solution is to be repeated for refrigerant-134a, refrigerant-22, and ammonia.

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Thermodynamics

-Thermodynamics is the study of the relationship between the energy

transformation in the system and other physical quantities such as temperature, pressure and volume (P, V, T). -A

thermodynamic equation of state is a mathematical

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relationship of the thermodynamic or state variables, such as pressure, volume and temperature.

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transfer is energy in

transition due to a

temperature

difference. The three

modes of heat

transfer are

conduction,

convection, and

radiation. Conduction

through Plane Walls

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Conduction heat transfer is a progressive exchange of energy between the molecules of a substance. Fourier's law of heat conduction is $Q = -kA \frac{dT}{dx}$ here Q!

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