

# Chapter 9 Stoichiometry Worksheet

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CH Ideal Stoichiometric Calculations Chapter 9 2 Mr C ~~Chapter 9 Stoichiometry Worksheet~~

9.1. Objectives. • Define stoichiometry. • Describe the importance of the mole ratio in stoichiometric calculations. • Write a mole ratio relating two substances in a chemical equation. 2/26/2018 2. Stoich Definitions. • Composition stoichiometry deals with the mass relationships of elements in compounds. • Reaction stoichiometry involves the mass relationships between reactants and products in a chemical reaction.

~~Chapter 9~~

stoichiometry (which you studied in Chapter 3) deals with the mass relationships of elements in compounds. Reaction stoichiometry involves the mass relationships between reactants and products in a chemical reaction. Reaction stoichiometry is

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the subject of this chapter and it is based on

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Stoichiometry Chapter 9, p. 275 - 294 Intro to Stoichiometry • Reaction Stoichiometry: Involves the mass relationships between reactants and products in a chemical reaction • Coefficients in a chemical reaction represent the mole ratios of each substance that react together • Example:  $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$

### ~~Chapter 9 Stoichiometry Notes~~

9.3 Objectives Describe a method for determining which of two reactants is a limiting reactant. Calculate the amount in moles or mass in grams of a product, given the amounts in moles or masses in grams of two reactants, one of which is in excess.; Distinguish between theoretical yield, actual yield, and percentage yield.; Calculate percentage yield, given the actual yield and

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Chapter 9 Worksheet Stoichiometry 1. Consider the following equation:  $\text{As}_2\text{O}_3 + 2\text{Cl}_2 + 5\text{H}_2\text{O} \rightarrow 2\text{H}_3\text{AsO}_4 + 4\text{HCl}$  a. How many moles of  $\text{H}_3\text{AsO}_4$  can be produced from 3.37 mol of  $\text{H}_2\text{O}$ ? b. How many grams of  $\text{H}_2\text{O}$  are needed to react with 0.789 mol of  $\text{Cl}_2$ ? c. How many moles of  $\text{As}_2\text{O}_3$  are needed to produce 62.7g of  $\text{HCl}$ ? d.

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CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of  $\text{N}_2$  are mixed with 12.0 mol of  $\text{H}_2$  according to the following equation:  $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$

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Reaction stoichiometry uses molar relationships to determine the amounts of unknown reactants or products from the amounts of known reactants or products. CHAPTER 9 DO NOT EDIT--Changes must be made through " File info " CorrectionKey=NL-A

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