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Ionic Bonding Introduction

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pH, pOH,  $\text{H}_3\text{O}^+$ ,  $\text{OH}^-$ ,  $K_w$ ,  $K_a$ ,  $K_b$ , pKa,  
and pKb Basic Calculations -Acids and  
Bases Chemistry Problems

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An ion consisting of a single atom is a monoatomic ion; an ion consisting of two or more atoms is referred to as a polyatomic ion. The positive electric charge of a proton is equal in magnitude to the negative charge of an electron; therefore, the net electric

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charge of an ion is equal to its number of protons minus its number of electrons.

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a concept of bonding theory based on the  
assumption that atoms tend to have either

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empty valence shells or full valence shells of eight electrons. ion. an atom of an element that has gained or lost one or more electrons, gaining a negative or positive charge in the process ... becoming ions in the process. transfer of electrons.



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Questions. Question ...

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Ions which are a part of the science subject  
Chemistry forms from atoms and electrons  
that have either gained or lost their weight by  
the removal or adding of one or more

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Answers  
valence electrons which would create either positive or a negative charge. The ions with a negative charge are called anions and the ones with a positive charge are called cations.

---

Cations and Anions - Difference Between



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Concept Review continued In the space  
provided at the left of each word or phrase,  
write the letter of the expression on the right

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that is most closely related. \_\_\_\_\_ 7.

hypothesis a. a well-tested explanation for a phenomenon based on observation, experimentation, and reasoning \_\_\_\_\_ 8.

theory b. a reasonable and testable explanation of ...

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## Skills Worksheet Concept Review

Why is an ionic compound unlikely to consist of two positively charged ions?

A3.1.1. Positive charges repel each other, so an ionic compound is not likely between two positively charged ions. H3.1.1. Think of Coulombs law. S3.1.1. Coulomb's law relates the force of interaction between two

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charged species.

$$F = \frac{Q_1 Q_2}{4\pi\epsilon_0 R_{12}^2}$$

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## 3.E: Ionic Bonding and Simple Ionic Compounds (Exercises I ...

1. All ionic compounds are solid at room

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Answers  
temperature. 2. Ionic compounds are hard and brittle - they do not break easily but when they do they fracture into crystals. 3. Ionic compounds have high melting and high boiling points. 4. Ionic compounds are good conductors of electric current in the liquid state. 5.

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Holt Chemistry Chapter 5 Ions and Ionic  
Compounds ...

The formation of positive and negative ions  
is discussed. Furthermore, the type of ion an  
element tends to form is a function of its  
position in the periodic table. Duration:  
9:45.

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15.2 The Formation of Ions | Conceptual  
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Chapter 5 Concepts review sheet:

5.1-kinetic and potential energy-work and  
heat as methods of energy transfer-system  
and surroundings 5.2-1 st law of



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Answers  
Thermodynamics-state function – internal energy  
5.3-state function – enthalpy  
5.4-enthalpy of reaction-remember that this is dependent upon amount of stuff reacting  
5.5-calorimetry – experimental method used to measure heat transfer in ...

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Chapter 5 review sheet - Chapter 5

Concepts review sheet 5 ...

This section is an important one. It provides more detailed information about ions than that found in Chapter 3, including how to predict their charges, how to convert between their names and symbols, and how they combine to form ionic compounds.

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Polyatomic ions (charged collections of atoms held together by covalent bonds) are also described.

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Chapter 5 Chemical Compounds

Concept Review j j. Section: What is

Ch € iSti ' Y Cornpiete each statement

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Answers  
below by underlining the correct word or phrase in brackets. 1. A chemical is any substance that has [de fi nit,e, indefinite} composition. Changes in chemicals, or chemical reactions, take place [only in test tubes, all. 2.

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